UPSC GEOLOGIST SYLLABUS

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APPENDIX – I
PLAN OF EXAMINATION
Plan of Examination
The Examination shall be conducted according to the following plan:—
(i) Stage-I : Combined Geo-Scientist (Preliminary) Examination (Computer Based Objective Type Papers) for the selection of candidates for the Stage-II:
........ 400 Marks
(ii) Stage-II : Combined Geo-Scientist (Main) Examination (Descriptive Type Papers); and
........ 600 Marks
(iii) Stage-III : Personality Test
........ 200 Marks
2. The candidates will first take the Computer Based Combined Geo-Scientist (Preliminary/Stage-I) Examination which consists of two Objective Type (multiple choices) questions papers for each stream. The Question Papers will be set in English only. The detailed Scheme of Stage-I is as follow:-
Stage-I: Combined Geo-Scientist (Preliminary) Examination:-

<table>
<thead>
<tr>
<th>Stream-I : Geologist &amp; Jr. Hydrogeologist</th>
</tr>
</thead>
<tbody>
<tr>
<td><strong>Subject</strong></td>
</tr>
<tr>
<td>------------------------------------------</td>
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<tr>
<td></td>
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</tbody>
</table>
Note-1: There will be penalty (Negative Marking) for wrong answers marked by a candidate in the objective type question papers.
(i) There are four alternatives for the answers to every question. For each question for which a wrong answer has been given by the candidate, one-third of the marks assigned to that question will be deducted as penalty.
(ii) If a question is left blank i.e. no answer is given by the candidate, there will be no penalty for that question.
Note-2: The candidates are not permitted to use calculators for answering Objective Type Papers. They should therefore not bring the same inside the Examination Hall.
Note-3: Only those candidates who are declared by the Commission to have qualified in the Preliminary/Stage-I Examination in the year will be eligible for admission to the Main/Stage-II Examination of that year provided they are otherwise eligible for admission to the Main/Stage-II Examination. The number of candidates to be admitted to the Main/Stage-II Examination will be about six to seven times the total approximate number of vacancies to be filled in the year through this Examination.
Note-4: The Commission will draw a list of candidates to be qualified for Combined Geo-Scientist (Main) Examination based on the criterion of minimum qualifying marks in General Studies Paper (Paper-I) and Geo-Scientist Stream specific paper (Paper-II) of Preliminary Examination.
3. The Combined Geo-Scientist (Main) Examination will consist of three conventional type papers for each stream. Conventional Type papers must be answered in English only. Question paper will be set in English only. The detailed scheme of Stage-II is as follows:-

Stream-I : Geologist

<table>
<thead>
<tr>
<th>Subject</th>
<th>Duration</th>
<th>Maximum Marks</th>
</tr>
</thead>
<tbody>
<tr>
<td>Paper-I : Geology</td>
<td>3 Hours</td>
<td>200 Marks</td>
</tr>
<tr>
<td>Paper-II : Geology</td>
<td>3 Hours</td>
<td>200 Marks</td>
</tr>
</tbody>
</table>
Paper-III : Geology | 3 Hours | 200 Marks
| Total | 600 Marks

### Stream-II : Geophysicist

<table>
<thead>
<tr>
<th>Subject</th>
<th>Duration</th>
<th>Maximum Marks</th>
</tr>
</thead>
<tbody>
<tr>
<td>Paper-I : Geophysics</td>
<td>3 Hours</td>
<td>200 Marks</td>
</tr>
<tr>
<td>Paper-II : Geophysics</td>
<td>3 Hours</td>
<td>200 Marks</td>
</tr>
<tr>
<td>Paper-III : Geophysics</td>
<td>3 Hours</td>
<td>200 Marks</td>
</tr>
<tr>
<td>Total</td>
<td></td>
<td>600 Marks</td>
</tr>
</tbody>
</table>

### Stream-III : Chemist

<table>
<thead>
<tr>
<th>Subject</th>
<th>Duration</th>
<th>Maximum Marks</th>
</tr>
</thead>
<tbody>
<tr>
<td>Paper-I : Chemistry</td>
<td>3 Hours</td>
<td>200 Marks</td>
</tr>
<tr>
<td>Paper-II : Chemistry</td>
<td>3 Hours</td>
<td>200 Marks</td>
</tr>
<tr>
<td>Paper-III : Chemistry</td>
<td>3 Hours</td>
<td>200 Marks</td>
</tr>
<tr>
<td>Total</td>
<td></td>
<td>600 Marks</td>
</tr>
</tbody>
</table>

### Stream-IV : Jr. Hydrogeologist

<table>
<thead>
<tr>
<th>Subject</th>
<th>Duration</th>
<th>Maximum Marks</th>
</tr>
</thead>
<tbody>
<tr>
<td>Paper-I : Geology</td>
<td>3 Hours</td>
<td>200 Marks</td>
</tr>
<tr>
<td>Paper-II : Geology</td>
<td>3 Hours</td>
<td>200 Marks</td>
</tr>
<tr>
<td>Paper-III : Hydrogeology</td>
<td>3 Hours</td>
<td>200 Marks</td>
</tr>
<tr>
<td>Total</td>
<td></td>
<td>600 Marks</td>
</tr>
</tbody>
</table>

Note-I: Candidates competing for selection to the posts of Geologist, Geophysicist, Chemist and Junior Hydrogeologist will be required to appear in all the subjects mentioned against respective category above.

Note-II: Candidates competing for selection for both the posts of Geologist and Jr. Hydrogeologist will be required to appear in all the subjects mentioned against Stream-I and IV above.

Note-III: If any candidate failed to appear in any one or more of above papers, meant for Main Examination for selection to the post of Geologist, Geophysicist, Chemist and Hydrogeologist, their candidature shall stand rejected and part of Main Examination attempted by him/her shall not be evaluated and counted for any purpose.

Note-IV: Credit will be given for orderly effective and exact expression combined with due economy of words in all subjects of Examination.

Note V: Candidates should use only international form of Indian numerals (e.g. 1, 2, 3, 4, 5 etc.) while answering question papers.

Note VI: Candidates will be allowed the use of Non-Programmable type Pocket Calculators in Descriptive Type Papers of the Examination. Programmable type calculators will not be allowed and the use of such calculators shall tantamount to resorting to unfair means by
the candidates. Loaning or interchanging of calculators in the Examination Hall is not permitted.

Note VII: Candidates should note that if any irrelevant matters/signages/marks etc. are found written in the answer script(s), which would not be related to any question/answer and/or would be having the potential to disclose the candidate’s identity, the Commission will impose a penalty of deduction of marks from the total marks from the total marks otherwise accruing to the candidate or will not be evaluate the said script(s) on this account.

4. Common instructions for Stage-I and Stage-II.

4.1 Candidates must mark/write the papers in their own hand. In no circumstances will they be allowed the help of a scribe to mark/write the answers for them. The Persons with Benchmark Disabilities in the categories of blindness, locomotor disability (both arm affected – BA) and cerebral palsy will be provided the facility of scribe, if desired by the person. In case of other category of Persons with Benchmark Disabilities as defined under section 2(r) of the RPWD Act, 2016, the facility of scribe will be allowed to such candidates on production of a certificate to the effect that the person concerned has physical limitation to write, and scribe is essential to write Examination on behalf, from the Chief Medical Officer/Civil Surgeon/Medical Superintendent of a Government Health Care institution as per proforma at Appendix–V.

4.2 The candidates have discretion of opting for his/her own scribe or request the Commission for the same. The details of scribe i.e. whether own or the Commission’s and the details of scribe in case candidates are bringing their own scribe, will be sought at the time of filling up the application form online as per proforma at Appendix-VI.

4.3 The qualification of the Commission’s scribe as well as own scribe will not be more than the minimum qualification criteria of the Examination. However, the qualification of the scribe should always be matriculate or above.

4.4 The Persons with Benchmark Disabilities in the category of blindness, locomotor disability (both arm affected – BA) and cerebral palsy will be allowed Compensatory Time of twenty minutes per hour of the Examination. In case of other categories of Persons with Benchmark Disabilities, this facility will be provided on production of a certificate to the effect that the person concerned has physical limitation to write from the Chief Medical Officer/Civil Surgeon/Medical Superintendent of a Government Health Care institution as per proforma at Appendix – V.

Note (1) : The eligibility conditions of a scribe, his/her conduct inside the Examination Hall and the manner in which and extent to which he/she can help the PwBD candidate in writing the Examination shall be governed by the instructions issued by the UPSC in this regard. Violation of all or any of the said instructions shall entail the cancellation of the candidature of the PwBD candidate in addition to any other action that the UPSC may take against the scribe.

Note (2) : The criteria for determining the percentage of visual impairment shall be as follows:

<table>
<thead>
<tr>
<th>Better eye Best Corrected</th>
<th>Worse eye Best Corrected</th>
<th>Per Cent Impairment</th>
<th>Disability category</th>
</tr>
</thead>
<tbody>
<tr>
<td>6/6 to 6/18</td>
<td>6/6 to 6/18</td>
<td>0%</td>
<td>0</td>
</tr>
<tr>
<td>6/24 to 6/60</td>
<td>6/6 to 6/18</td>
<td>10%</td>
<td>0</td>
</tr>
<tr>
<td>Less than 6/60 to 3/60</td>
<td>Less than 6/60 to 3/60</td>
<td>20%</td>
<td>I</td>
</tr>
<tr>
<td>Less than 3/60 to No Light Perception</td>
<td>30%</td>
<td>II (One eyed person)</td>
<td></td>
</tr>
<tr>
<td>6/24 to 6/60</td>
<td>6/24 to 6/60</td>
<td>40%</td>
<td>III a (low vision)</td>
</tr>
<tr>
<td>Or</td>
<td>Less than 6/60 to 3/60</td>
<td>50%</td>
<td>III b (low vision)</td>
</tr>
<tr>
<td>Visual field less than 40 up to 20 degree around centre of fixation or heminaopia involving macula</td>
<td>Less than 3/60 to No Light Perception</td>
<td>60%</td>
<td>III c (low vision)</td>
</tr>
<tr>
<td>Less than 6/60 to 3/60</td>
<td>Less than 6/60 to 3/60</td>
<td>70%</td>
<td>III d (low vision)</td>
</tr>
<tr>
<td>Or</td>
<td>Less than 3/60 to No Light Perception</td>
<td>80%</td>
<td>III e (low vision)</td>
</tr>
<tr>
<td>Fixation of fixation</td>
<td>Less than 3/60 to No Light Perception</td>
<td>90%</td>
<td>IV a (Blindness)</td>
</tr>
<tr>
<td>----------------------</td>
<td>---------------------------------------</td>
<td>-----</td>
<td>-----------------</td>
</tr>
<tr>
<td>Less than 3/60 to 1/60 Or Visual field less than 10 degree around centre of fixation</td>
<td>Only HMCF Only Light Perception No Light Perception</td>
<td>100%</td>
<td>IV b (Blindness)</td>
</tr>
</tbody>
</table>

The concession admissible to blind candidates shall not be admissible to those suffering from Myopia.

4.5. In the question papers, wherever necessary, questions involving the Metric System of Weights and Measures only will be set.

5. The Stage-III will consist of Personality Test carrying 200 Marks (with no minimum qualifying marks). Candidates, who obtain such minimum qualifying marks in Stage-II as may be fixed by the Commission as per its discretion, shall be summoned by them for Stage-III (Personality Test). The number of candidates to be summoned for Personality Test will be about twice the number of vacancies to be filled. In the Personality Test, the candidate will be interviewed by Board i.e. Interview Board (s) constituted by the Commission. The object of the interview will be to assess the suitability for the post(s) of Geologist, Geophysicist, Chemist and Jr. Hydrogeologist. Special attention will be paid in the Personality Test to assessing the candidate’s capacity for leadership, initiative and intellectual curiosity, tact and other social qualities, mental and physical energy powers of practical application, integrity of character and aptitude for adapting themselves to the field life.

6. **The Details of the syllabi for Stage-I: (Preliminary Examination) and Stage-II: (Main Examination) of Combined Geo-Scientist Examination are as under:**

**SCHEDULE**

**STANDARD AND SYLLABUS**

Paper-I in General Studies of Stage-I is common for all streams and its standard will be such as may be expected of a science graduate. Paper-II of Stage-I (Stream specific) and 3 compulsory papers of Stage-II each on Geology, Geophysics, Chemistry and Hydrogeology subjects will be approximately of the M.Sc. degree standard of an Indian University and questions will generally be set to test the candidate’s grasp of the fundamentals in each subject.

There will be no practical examination in any of the subjects

**Syllabus of Combined Geo-Scientist (Preliminary) Examination**

**Stage-I (Objective Type)**

**Paper-I : General Studies (Common for all streams)**

- Current events of national and international importance.
- History of India and Indian National Movement.
- Indian and World Geography -Physical, Social, Economic Geography of India and the World.
- Indian Polity and Governance -Constitution, Political System, Panchayati Raj, Public Policy, Rights Issues, etc.
- Economic and Social Development – Sustainable Development, Poverty, Inclusion, Demographics, Social Sector initiatives, etc.
- General issues on Environmental Ecology, Bio-diversity and Climate Change - that do not require subject specialisation
- General Science

**Stage-I (Objective Type)**

**Paper-II : Geology/Hydrogeology**

1. **Physical Geology**
   
   Principle of uniformitarianism; origin, differentiation and internal structure of the Earth; origin of atmosphere; earthquakes and volcanoes; continental drift, sea-floor spreading, isostasy, orogeny and plate tectonics; geological action of rivers, wind, glaciers, waves; erosional and depositional landforms; weathering processes and products.

2. **Structural Geology**
Stress, strain and rheological properties of rocks; planar and linear structures; classification of folds and faults; Mohr's circle and criteria for failure of rocks; ductile and brittle shear in rocks; study of toposheets, V-rules and outcrop patterns; stereographic projections of structural elements.

3. **Mineralogy**
   Elements of symmetry, notations and indices; Bravais lattices; chemical classification of minerals; isomorphism, polymorphism, solid solution and exsolution; silicate structures; physical and optical properties of common rock forming minerals- olivine, garnet, pyroxene, amphibole, mica, feldspar and quartz.

4. **Igneous Petrology**
   Magma types and their evolution; IUGS classification of igneous rocks; forms, structures and textures of igneous rocks; applications of binary and ternary phase diagrams in petrogenesis; magmatic differentiation and assimilation; petrogenesis of granites, basalts, komatiites and alkaline rocks (carbonatite, kimberlite, lamprophyre and nepheline syenite).

5. **Metamorphic Petrology**
   Limits, types and controls of metamorphism; metamorphic structures- slate, schist and gneiss; metamorphic textures- pre, syn and post tectonic porphyroblasts; concept of metamorphic zone, isograd and facies; geothermal gradients, facies series and plate tectonics.

6. **Sedimentology**
   Origin of sediments; sedimentary textures, grain-size scale; primary sedimentary structures; classification of sandstone and carbonate rocks; siliciclastic depositional environments and sedimentary facies; diagenesis of carbonate sediments.

7. **Paleontology**
   Fossils and processes of fossilization; concept of species and binomial nomenclature; morphology and classification of invertebrates (Trilobites, Brachiopods, Lamellibranchs, Gastropods and Cephalopods); evolution in Equidae and Hominidae; microfossils-Foraminifera, Ostracoda; Gondwana flora.

8. **Stratigraphy**
   Law of superposition; stratigraphic nomenclature- lithostratigraphy, biostratigraphy and chronostratigraphy; Archaean cratonic nuclei of Peninsular India (Dharwar, Singhbum, and Aravalli cratons); Proterozoic mobile belts (Central Indian Tectonic Zone, Aravalli-Delhi and Eastern Ghat); Purana sedimentary basins (Cuddapah and Vindhyan); Phanerozoic stratigraphy of India- Spiti, Kashmir, Damodar valley, Kutch, Trichinopoly, Siwaliks and Indo-Gangetic alluvium.

9. **Economic Geology**
   Properties of mineral deposits- form, mineral assemblage, texture, rock-ore association and relationship; magmatic, sedimentary, metamorphic, hydrothermal, supergene and weathering-related processes of ore formation; processes of formation of coal and petroleum; distribution and geological characteristics of major mineral and hydrocarbon deposits of India.

10. **Hydrogeology**
    Groundwater occurrence and aquifer characteristics, porosity, permeability, hydraulic conductivity, transmissivity; Darcy's Law in homogenous and heterogenous media; Bernoulli equation, Reynold's number; composition of groundwater; application of H and O isotopes in groundwater studies; artificial recharge of groundwater.

**Stage-I (Objective Type)**

**Paper-II : Geophysics**

1. **Solid Earth Geophysics:**

2. **Mathematical Methods in Geophysics:**
   Elements of vector analysis, Vector algebra, Properties of scalars, vectors and tensors, Gradient, Divergence and Curl, Gauss's divergence theorem, Stoke's theorem. Matrices, Eigen values


3. **Electromagnetism:**
Electrostatic and magneto-static fields, Coulomb's law, Electrical permittivity and dielectric constant, Lorentz force and their applications. Ampere’s law, Biot and Savart’s law, Gauss’s Theorem, Poisson’s equation. Laplace’s equation: solution of Laplace’s equation in Cartesian coordinates, use of Laplace’s equation in the solutions of geophysical and electrostatic problems. Displacement current, Faraday’s law of electromagnetic induction. Maxwell’s equations. Boundary conditions. Wave equation, plane electromagnetic waves in free space, dielectric and conducting media, electromagnetic vector and scalar potentials.

4. **Geophysical Prospecting:**

5. **Remote Sensing and Thermodynamics:**

6. **Nuclear Physics and Radiometry:**

**Stage-I (Objective Type)**

**Paper-II : Chemistry**

1. **Chemical periodicity:**
Schrödinger equation for the H-atom. Radial distribution curves for 1s, 2s, 2p, 3s, 3p, 3d orbitals. Electronic configurations of multi-electron atoms.

Periodic table, group trends and periodic trends in physical properties. Classification of elements on the basis of electronic configuration. Modern IUPAC Periodic table. General characteristics of s, p, d and f block elements. Effective nuclear charges, screening effects, atomic radii, ionic radii, covalent radii. Ionization enthalpy, electron gain enthalpy and electronegativity. Group trends and periodic trends in these properties in respect of s-, p- and d-block elements. General trends of variation of electronic configuration, elemental forms, metallic nature, magnetic properties, catenation and catalytic properties, oxidation states, aqueous and redox chemistry in common oxidation states, properties and reactions of important compounds such as hydrides, halides, oxides, oxy-acids, complex chemistry in respect of s-block and p-block elements.
2. **Chemical bonding and structure:**

   **Ionic bonding:** Size effects, radius ratio rules and their limitations. Packing of ions in crystals, lattice energy, Born-Landé equation and its applications, Born-Haber cycle and its applications. Solvation energy, polarizing power and polarizability, ionic potential, Fajan’s rules. Defects in solids.

   **Covalent bonding:** Valence Bond Theory, Molecular Orbital Theory, hybridization. Concept of resonance, resonance energy, resonance structures.

   **Coordinate bonding:** Werner theory of coordination compounds, double salts and complex salts. Ambidentate and polydentate ligands, chelate complexes. IUPAC nomenclature of coordination compounds. Coordination numbers, Geometrical isomerism. Stereoisomerism in square planar and octahedral complexes.

3. **Acids and bases:**

   Chemical and ionic equilibrium. Strengths of acids and bases. Ionization of weak acids and bases in aqueous solutions, application of Ostwald’s dilution law, ionization constants, ionic product of water, pH-scale, effect of temperature on pH, buffer solutions and their pH values, buffer action & buffer capacity; different types of buffers and Henderson’s equation.

4. **Theoretical basis of quantitative inorganic analysis:**

   **Volumetric Analysis:** Equivalent weights, different types of solutions, normal and molar solutions. Primary and secondary standard substances. General principles of different types of titrations: i) acid-base, ii) redox, iii) complexometric, iv) Precipitation. Types of indicators - i) acid-base, ii) redox iii) metal-ion indicators.

5. **Kinetic theory and the gaseous state:**

   Kinetic theory of gases, average kinetic energy of translation, Boltzmann constant and absolute scale of temperature. Maxwell-Boltzmann distribution of speeds. Calculations of average, root mean square and most probable velocities. Collision diameter; collision number and mean free path; frequency of binary collisions; wall collision and rate of effusion.

6. **Chemical thermodynamics and chemical equilibrium:**

   First law and its applications to chemical problems. Thermodynamic functions. Total differentials and state functions. Free expansion, Joule-Thomson coefficient and inversion temperature. Hess’ law. Applications of Second law of thermodynamics. Gibbs function (G) and Helmholtz function (A), Gibbs-Helmholtz equation, criteria for thermodynamic equilibrium and spontaneity of chemical processes.

7. **Solutions of non-electrolytes:**

   Colligative properties of solutions, Raoult’s Law, relative lowering of vapour pressure, osmosis and osmotic pressure; elevation of boiling point and depression of freezing point of solvents. Solubility of gases in liquids and solid solutions.

8. **Electrochemistry:**


9. **Basic organic chemistry:**

   Delocalized chemical bond, resonance, conjugation, hyperconjugation, hybridisation, orbital pictures of bonding sp³, sp², sp: C-C, C-N and C-O system), bond polarization and bond polarizability. Reactive intermediates: General methods of formation, relative stability and reactivity of carbocations, carbanions and free radicals.

10. **Stereochemistry:**

    Configuration and chirality (simple treatment of elements of symmetry), optical isomerism of compounds containing two to three stereogenic centres, R,S nomenclature, geometrical isomerism in compounds containing two C=C double bonds (E,Z naming), and simple cyclic systems, Newman projection (ethane and substituted ethane).

11. **Types of organic reactions:**
Aliphatic substitution reactions: $S_{N1}$, $S_{N2}$ mechanisms, stereochemistry, relative reactivity in aliphatic substitutions. Effect of substrate structure, attacking nucleophile, leaving group and reaction medium and competitive reactions.

Elimination reactions: $E_1$, $E_2$, mechanisms, stereochemistry, relative reactivity in aliphatic eliminations. Effect of substrate structure, attacking base, leaving group, reaction medium and competitive reactions, orientation of the double bond, Saytzeff and Hoffman rules.

Addition reactions: Electrophilic, nucleophilic and radical addition reactions at carbon-carbon double bonds.

Electrophilic and nucleophilic aromatic substitution: Electrophilic (halogenation, sulphonation, nitration, Friedal-Crafts alkylation and acylation), nucleophilic (simple $S_{NAr}$, $S_{N1}$ and aryne reactions).

12. Molecular Rearrangements:
   Acid induced rearrangement and Wagner-Meerwein rearrangements. Neighbouring group participation.

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Syllabus of Combined Geo-Scientist (Main) Examination
Stage-II (Descriptive Type)
Geology: Paper-I

Section A. Physical geology and remote sensing
Evolution of Earth; Earth's internal structure; earthquakes and volcanoes; principles of geodesy, isostasy; weathering - processes and products; geomorphic landforms formed by action of rivers, wind, glaciers, waves and groundwater; features of ocean floor; continental shelf, slope and rise; concepts of landscape evolution; major geomorphic features of India - coastal, peninsular and extrapeninsular.
Electromagnetic spectrum; electromagnetic bands in remote sensing; spectral signatures of soil, rock, water and vegetation; thermal, near infra-red and microwave remote sensing; digital image processing; LANDSAT, IRS and SPOT- characteristics and use; aerial photos- types, scale, parallax, relief displacement; elements of image interpretation.

Section B. Structural geology
Principles of geological mapping; kinematic and dynamic analysis of deformation; stress-strain relationships for elastic, plastic and viscous materials; measurement of strain in deformed rocks; structural analysis of fold, cleavage, boudin, lineation, joint, and fault; stereographic projection of linear and planar structures; superposed deformation; deformation at microscale- dynamic and static recrystallisation, controls of strain rate and temperature on development of microfabrics; brittle and ductile shear zones; time relationship between crystallisation and deformation, calculation of paleostress.

Section C. Sedimentology
Classification of sedimentary rocks; sedimentary textures-grain size, roundness, sphericity, shape and fabric; quantitative grain size analysis; sediment transport and deposition- fluid and sediment gravity flows, laminar and turbulent flows, Reynold's number, Froude number, grain entrainment, Hjulstrom diagram, bed load and suspension load transport; primary sedimentary structures; penecontemporaneous deformation structure; biogenic structures; principles and application of paleocurrent analysis; composition and significance of different types of sandstone, limestone, banded iron formation, mudstone, conglomerate; carbonate diagenesis and dolomitisation; sedimentary environments and facies-facies models for fluvial, glacial, deltaic, siliciclastic shallow and deep marine environments; carbonate platforms- types and facies models; sedimentation in major tectonic settings; principles of sequence stratigraphy-concepts and factors controlling base level changes, parasequence, clinoform, systems tract, unconformity and sequence boundary.

Section D. Paleontology
Fossil record and geological time scale; modes of preservation of fossils and concept of taphonomy; body- and ichno-fossils, species concept, organic evolution, Ediacara Fauna; morphology and time range of Graftolites, Trilobites, Brachiopods, Lamellibranchs, Gastropods, Cephalopods, Echinoids and Corals; evolutionary trends in Trilobites, Lamellibranchs, Gastropods and Cephalopods; micropaleontology- methods of preparation of microfossils, morphology of microfossil groups
(Foraminifera, Ostracoda), fossil spores, pollen and dinoflagellates; Gondwana plant fossils and their significance; vertebrate life through ages, evolution in Proboscidea, Equidae and Hominidae; applications of paleontological data in stratigraphy, paleoecology and paleoclimatology; mass extinctions.

**Section E. Stratigraphy**

Principles of stratigraphy-code of stratigraphic nomenclature of India; lithostratigraphy, biostratigraphy, chronostratigraphy and magnetostratigraphy; principles of stratigraphic correlation; characteristics of Archean granite-greenstone belts; Indian stratigraphy- geological evolution of Archean nuclei (Dharwar, Bastar, Singhbhum, Aravalli and Bundelkhand); Proterozoic mobile belts-Eastern Ghats Mobile Belt, Southern Granulite Terrain, Central Indian Tectonic Zone, Aravalli-Delhi Belt, North Singhbhum Mobile Belt; Proterozoic sedimentary basins (Cuddapah and Vindhyan); Phanerozoic stratigraphy- Paleozoic (Spiti, Kashmir and Kumaon), Mesozoic (Spiti, Kutch, Narmada Valley and Trichinopoly), Gondwana Supergroup, Cenozoic (Assam, Bengal basins, Garhwal-Shimla Himalayas); Siwaliks; boundary problems in Indian stratigraphy.
slopes of metamorphic reactions; heat flow, diffusion and mass transfer; Fourier’s law of heat conduction; geothermobarometry; mass and energy change during fluid-rock interactions; charnockite problem, formation of skarns, progressive and retrogressive metamorphism of pelitic, calcareous and basic rocks; P-T-t path and tectonic setting.

**Section E. Geodynamics**

Phase transitions and seismic discontinuities in the Earth; seismic waves and relation between V_p, V_s and density; seismic and petrological Moho; rheology of rocks and fluids (Newtonian and non-Newtonian liquids); rock magnetism and its origin; polarity reversals, polar wandering and supercontinent cycles; continental drift, sea floor spreading; gravity and magnetic anomalies of ocean floors and their significance; mantle plumes and their origin; plate tectonics- types of plate boundaries and their inter-relationship; heat flow and heat production of the crust.

**Stage-II (Descriptive Type)**

**Geology : Paper-III**

**Section A. Economic geology**

Ore minerals and industrial minerals; physical and optical properties of ore minerals; ore textures and paragenesis; characteristics of mineral deposits- spatial and temporal distribution, rock-ore association; syngenetic and epigenetic deposits, forms of ore bodies, stratiform and strata-bound deposits; ore forming processes- source and migration of ore constituents and ore fluid, mechanism of ore deposition; magmatic and pegmatitic deposits (chromite, Ti-magnetite, diamond, Cu-Ni sulphide, PGE, REE, muscovite, rare metals); hydrothermal deposits (porphyry Cu-Mo, greisen Sn-W, skarn, VMS and SEDEX type sulphide deposits, orogenic gold); sedimentary deposits (Fe, Mn, phosphorite, placer); supergene deposits (Cu, Al, Ni and Fe); metamorphic and metamorphosed deposits (Mn, graphite); fluid inclusions in ore mineral assemblage- physical and chemical properties, microthermometry; stable isotope (S, C, O, H) in ore genesis- geothermometry, source of ore constituents; global tectonics and mineralisation.

**Section B. Indian mineral deposits and mineral economics**

Distribution of mineral deposits in Indian shield; geological characteristics of important industrial mineral and ore deposits in India- chromite, diamond, muscovite, Cu-Pb-Zn, Sn-W, Au, Fe-Mn, bauxite; minerals used in refractory, fertilizer, ceramic, cement, glass, paint industries; minerals used as abrasive, filler; building stones.

Strategic, critical and essential minerals; India’s status in mineral production; co-products and by-products; consumption, substitution and conservation of minerals; National Mineral Policy; Mineral Concession Rules; marine mineral resources and laws of the sea.

**Section C. Mineral exploration**

Stages of exploration; scope, objectives and methods of prospecting, regional exploration and detailed exploration; geological, geochemical and geobotanical methods; litho-, bio-, soil geochemical surveys, mobility and dispersion of elements, geochemical anomalies; ore controls and guides; pitting, trenching, drilling; sampling, assaying, ore reserve estimation; categorization of ore reserves; geophysical methods- ground and airborne surveys; gravity, magnetic, electrical and seismic methods of mineral exploration.

**Section D. Fuel geology and Engineering geology**

Coal and its properties; proximate and ultimate analysis; different varieties and ranks of coal; concept of coal maturity, peat, lignite, bituminous and anthracite coal; origin of coal, coalification process; lithotypes, microlithotypes and maceral groups of coal; mineral and organic matter in coal; lignite and coal deposits of India; origin, migration and entrapment of natural hydrocarbons; characteristics of source and reservoir rocks; structural, stratigraphic and mixed traps; geological, geochemical and geophysical methods of hydrocarbon exploration; petroliferous basins of India; geological characteristics and genesis of major types of U deposits and their distribution in India.

Engineering properties of rocks; geological investigations in construction of dams, reservoirs, tunnels, bridges, highways and coastal protection structures; geologic considerations of construction materials.

**Section E. Environmental geology and Natural hazards**

Stefan-Boltzmann equation and planetary temperature; cause and effects of global climate change; Earth’s radiation budget; greenhouse gases and effect; examples of positive and negative feedback
mechanisms; biogeochemical cycle of carbon; geological investigations of nuclear waste disposal sites; marginal marine environments- estuaries, mangroves and lagoons; ozone hole depletion, ocean acidification, coral bleaching, Milankovitch cycle, sea level rise, eutrophication and acid rain; environmental impacts of urbanization, mining and hydropower projects; water pollution, water logging and soil erosion; Himalayan glaciers; causes and consequences of earthquakes, volcanoes, tsunami, floods, landslides, coastal erosion, droughts and desertification; application of remote sensing and geographic information systems (GIS) in environmental management.

**Stage-II (Descriptive Type)**

**Hydrogeology**

**Section A. Occurrence and distribution of groundwater**

Origin of water on Earth; global water cycle and budget; residence time concept, geologic formations as aquifers; confined and unconfined aquifers; groundwater table mapping and piezometric nests; porosity, void ratio, effective porosity and representative porosity range; primary and secondary porosities; groundwater zonation; specific retention, specific yield; groundwater basins; springs.

**Section B. Groundwater movement and well hydraulics**

Groundwater flow concepts; Darcy’s Law in isotropic and anisotropic media and validity; water flow rates, direction and water volume in aquifers; permeability and hydraulic conductivity and ranges in representative rocks; Bernoulli equation; determination of hydraulic conductivity in field and laboratory; concept of groundwater flow through dispersion and diffusion; transmissivity and aquifer thickness.

**Section C. Water wells and groundwater levels**

Unidirectional and radial flow to a well (steady and unsteady); well flow near aquifer boundaries; methods for constructing shallow wells, drilling wells, well completion; testing wells, pumping test, slug tests for confined and unconfined aquifers; fluctuations in groundwater levels; stream flow and groundwater flows; groundwater level fluctuations; land subsidence; impact of global climate change on groundwater.

**Section D. Groundwater exploration**

Surface investigation of groundwater- geologic, remote sensing, electrical resistivity, seismic, gravity and magnetic methods; sub-surface investigation of groundwater- test drilling, resistivity logging, spontaneous potential logging, radiation logging.

**Section E. Groundwater quality and management**

Groundwater composition, units of expression, mass-balance calculations; rock-water interaction (chemical equilibrium, free energy, redox reactions and cation/anion exchanges), graphic representation of chemical data; groundwater hardness, microorganisms in groundwater; water quality standards; sea-water intrusion; groundwater issues due to urbanization; solid and liquid waste disposal and plume migration models; application of isotopes (H, C, O) in groundwater; concepts of artificial recharge methods; managing groundwater resources; groundwater basin investigations and management practices.

**Stage-II (Descriptive Type)**

**Geophysics : Paper-I**

**PART-A**

**A1. Solid Earth Geophysics:**


**A2. Earthquake Seismology:**

Seismology, earthquakes, focal depth, epicenter, great Indian earthquakes, Intensity and Magnitude scales, Energy of earthquakes, foreshocks, aftershocks, Elastic rebound theory, Types and Nature of faulting, Fault plane solutions, Seismicity and Seismotectonics of India, Frequency-Magnitude relation (b-values). Bulk and rigidity modulus, Lame’s Parameter, Seismic waves: types and their
propagation characteristics, absorption, attenuation and dispersion. Seismic ray theory for spherically and horizontally stratified earth, basic principles of Seismic Tomography and receiver function analysis, Velocity structure, Vp/Vs studies, Seismic network and arrays, telemetry systems, Principle of electromagnetic seismograph, displacement meters, velocity meters, accelerometers, Broadband Seismometer, WWSSN stations, seismic arrays for detection of nuclear explosions. Earthquake prediction; dilatancy theory, short-, medium- and long- term predictions, Seismic microzonations, Applications for engineering problems.

A3. Mathematical methods in Geophysics:
Elements of vector analysis, Gradient, Divergence and Curl, Gauss’s divergence theorem, Stoke’s theorem, Gravitational field, Newton’s Law of gravitation, Gravitation potential and fields due to bodies of different geometric shapes, Coulomb’s law, Electrical permittivity and dielectric constant, Origin of Magnetic field, Ampere’s law, Biot and Savart’s law, Geomagnetic fields, Magnetic fields due to different type of structures, Solution of Laplace equation in Cartesian, Cylindrical and Spherical Coordinates, Image theory, Electrical fields due to charge, point source, continuous charge distribution and double layers, equipotential and line of force. Current and potential in the earth, basic concept and equations of electromagnetic induction, Maxwell’s Equation, near and far fields, Attenuation of EM waves, EM field of a loops of wire on half space and multi-layered media.

A4. Geophysical Inversion:

PART-B:

B1. Mathematical Methods of Physics:
Dimensional analysis; Units and measurement; Vector algebra and vector calculus; Linear algebra, Matrices: Eigenvalues and eigenvectors; Linear ordinary differential equations of first and second order; Special functions (Hermite, Bessel, Laguerre and Legendre); Fourier series, Fourier and Laplace transforms; Elementary probability theory, Random variables, Binomial, Poisson and normal distributions; Green’s function; Partial differential equations (Laplace, wave and heat equations in two and three dimensions); Elements of numerical techniques: root of functions, interpolation, and extrapolation, integration by trapezoid and Simpson’s rule, solution of first order differential equation using Runge-Kutta method; Tensors; Complex variables and analysis; Analytic functions; Taylor & Laurent series; poles, residues and evaluation of integrals; Beta and Gamma functions. Operators and their properties; Least-squares fitting.

B2. Electrodynamics:
Electrostatics: Gauss’ Law and its applications; Laplace and Poisson equations, Boundary value problems; Magnetostatics: Biot-Savart law, Ampere’s theorem; Ampere’s circuital law; Magnetic vector potential; Faraday’s law of electromagnetic induction; Electromagnetic vector and scalar potentials; Uniqueness of electromagnetic potentials and concept of gauge: Lorentz and Coulomb gauges; Lorentz force; Charged particles in uniform and non-uniform electric and magnetic fields; Poynting theorem; Electromagnetic fields from Lienard-Wiechert potential of a moving charge; Bremsstrahlung radiation; Cerenkov radiation; Radiation due to oscillatory electric dipole; Condition for plasma existence; Occurrence of plasma; Magnetohydrodynamics; Plasma waves; Transformation of electromagnetic potentials; Lorentz condition; Invariance or covariance of Maxwell field equations in terms of 4 vectors; Electromagnetic field tensor; Lorentz transformation of electric and magnetic fields.
B3. Electromagnetic Theory:
Maxwell's equations: its differential and integral forms, physical significance; Displacement current; Boundary conditions; Wave equation, Plane electromagnetic waves in: free space, non-conducting isotropic medium, conducting medium; Scalar and vector potentials; Reflection; refraction of electromagnetic waves; Fresnel's Law; interference; coherence; diffraction and polarization; Lorentz invariance of Maxwell's equations; Transmission lines and waveguides.

B4. Introductory Atmospheric and Space Physics:
The neutral atmosphere; Atmospheric nomenclature; Height profile of atmosphere; Hydrostatic equation; Geopotential height; Expansion and contraction; Fundamental forces in the atmosphere; Apparent forces; Atmospheric composition; Solar radiation interaction with the neutral atmosphere; Climate change; Electromagnetic radiation and propagation of Waves: EM Radiation; Effects of environment; Antennas: basic considerations, types. Propagation of waves: ground wave, sky wave, and space wave propagation; troposcatter communication and extra terrestrial communication; The Ionosphere; Morphology of ionosphere: the D, E and F-regions; Chemistry of the ionosphere Ionospheric parameters E and F region anomalies and irregularities in the ionosphere; Global Positioning Systems (GPS): overview of GPS system, augmentation services GPS system segment; GPS signal characteristics; GPS errors; multi path effects; GPS performance; Satellite navigation system and applications.

Stage-II (Descriptive Type)
Geophysics : Paper-II

PART-A
A1. Potential Field (Gravity and Magnetic) Methods:

A2. Electrical and Electromagnetic methods:
Electrical properties of rocks and minerals, concepts and assumptions of horizontally stratified earth, anisotropy and its effects on electrical fields, geoelectric and geological sections, D.C Resistivity method. Concept of natural electric field, various electrode configurations, Profiling and Sounding (VES). Types of Sounding curves, Equivalence and Suppression, Concept of Electrical Resistivity Tomography (ERT). SP Method,: Origin of SP, application of SP surveys. Induced Polarization (IP) Method: Origin of IP, Membrane and Electrode polarization, time and frequency domains of measurement, chargeability, percent frequency effect and metal factor, Application of IP surveys for mineral exploration. Electromagnetic methods, Passive and Active source methods, Diffusion equation, wave equation and damped wave equation used in EM method, boundary conditions, skin depth, depth of investigation and depth of penetration, amplitude and phase relations, real and imaginary components, elliptical polarization, Principles of EM prospecting, various EM methods: Dip angle, Turam, moving source-receiver methods-horizontal loop (Slingram), AFMAG, and VLF., Principles of Time Domain EM: INPUT method. EM Profiling and sounding, Interpretation of EM anomalies. Principle of EM scale modeling. Magnetotelluric methods: Origin and characteristics of MT fields, Instrumentation, Transverse Electric and Transverse Magnetic Modes,

A3. Seismic Prospecting:
Basic principles of seismic methods, Various factors affecting seismic velocities in rocks, Reflection, refraction and Energy partitioning at an interface, Geometrical spreading, Reflection and refraction of wave phenomena in a layered and dipping media. Seismic absorption and anisotropy, Multi channel seismic (CDP) data acquisition (2D and 3D), sources of energy, Geophones, geometry of arrays, different spread geometry, Instrumentation, digital recording. Different types of multiples, Travel time curves, corrections, Interpretation of data, bright spot, low velocity layer, Data processing, static and dynamic (NMO and DMO) corrections, shot-receiver gather, foldage, multiplexing and demultiplexing. Dix's equation, Velocities: Interval, Average and RMS, Seismic resolution and Fresnel Zone, Velocity analysis and Migration techniques, Seismic Interpretation, Time and Depth Section, Fundamentals of VSP method, High Resolution Seismic Surveys (HRSS).

A4. Borehole Geophysics:
Objectives of well logging, concepts of borehole geophysics, borehole conditions, properties of reservoir rock formations, formation parameters and their relationships-formation factor, porosity, permeability, formation water resistivity, water saturation, irreducible water saturation, hydrocarbon saturation, residual hydrocarbon saturation; Arhcie’s and Humble’s equations; principles, instrumentations, operational procedures and interpretations of various geophysical logs: SP, resistivity and micro resistivity, gamma ray, neutron, sonic, temperature, caliper and directional logs. Production logging, overlay and cross-plots of well-log data, determination of formation lithology, porosity, permeability and oil-water saturation, sub-surface correlation and mapping, delineation of fractures; application of well-logging in hydrocarbon, groundwater, coal, metallic and non-metallic mineral exploration.

PART-B

B1. Classical Mechanics
Inertial and non-inertial frames, Newton’s laws; Pseudo forces; Central force motion; Two-body collisions, Scattering in laboratory and centre-of-mass frames; Rigid body dynamics, Moment of inertia, Variational principle, Lagrangian and Hamiltonian formalisms and equations of motion; Poisson brackets and canonical transformations; Symmetry, Invariance and conservation laws, Cyclic coordinates; Periodic motion, Small oscillations and normal modes; Special theory of relativity, Lorentz transformations, Relativistic kinematics and mass-energy equivalence.

B2. Thermodynamics and Statistical Physics
Laws of thermodynamics and their significance; Thermodynamic potentials, Maxwell relations; Chemical potential, Phase equilibria; Phase space, Micro- and macro- states; Micro canonical, canonical and grand-canonical ensembles and partition functions; Free Energy and connection with thermodynamic quantities; First and second order phase transitions; Maxwell-Boltzmann distribution, Quantum statistics, Ideal Fermi and Bose gases; Principle of detailed balance; Blackbody radiation and Planck’s distribution law; Bose-Einstein condensation; Random walk and Brownian motion; Diffusion equation.

B3. Atomic and Molecular Physics and Characterization of materials
Quantum states of an electron in an atom; Electron spin; Stern-Gerlach experiment; Spectrum of Hydrogen, Helium and alkali atoms; Relativistic corrections for energy levels of hydrogen; Hyperfine structure and isotopic shift; Width of spectral lines; LS and JJ coupling; Zeeman, Paschen Back and Stark effects; Rotational, vibrational, electronic, and Raman spectra of diatomic molecules; Frank-Condon principle; Thermal and optical properties of materials, Study of microstructure using SEM, Study of crystal structure using TEM, Resonance methods: Spin and applied magnetic field, Larmor precession, relaxation times - spin-spin relaxation, Spin-lattice relaxation, Electron spin resonance, g factor, Nuclear Magnetic resonance, line width, Motional narrowing, Hyperfine splitting; Nuclear Gamma Resonance: Principles of Mössbauer Spectroscopy, Line width, Resonance absorption, Isomer Shift, Quadrupole splitting.

B4. Nuclear and Particle Physics
Basic nuclear properties: size, shape, charge distribution, spin and parity; Binding energy, Packing fraction, Semi-empirical mass formula; Liquid drop model; Fission and fusion, Nuclear reactor; Line...
of stability, Characteristics of the nuclear forces, Nucleon-nucleon potential; Charge-independence and charge-symmetry of nuclear forces; Isospin; Deuteron problem; Evidence of shell structure, Single-particle shell model and, its validity and limitations; Elementary ideas of alpha, beta and gamma decays and their selection rules; Nuclear reactions, reaction mechanisms, compound nuclei and direct reactions; Classification of fundamental forces; Elementary particles (quarks, baryons, mesons, leptons); Spin and parity assignments, strangeness; Gell Mann-Nishijima formula; C, P and T invariance and applications of symmetry arguments to particle reactions, Parity non-conservation in weak interaction; Relativistic kinematics.

Stage-II (Descriptive Type)

Geophysics : Paper-III

PART-A

A1. Radiometric and Airborne Geophysics:
Principles of radioactivity, radioactivity decay processes, units, radioactivity of rocks and minerals, Instruments, Ionization chamber, G-M counter, Scintillation counter, Gamma ray spectrometer, Radiometric prospecting for mineral exploration (Direct/Indirect applications), beach placers, titanium, zirconium and rare-earths, radon studies in seismology and environmental applications. Airborne geophysical surveys (gravity, magnetic, electromagnetic and radiometric), planning of surveys, flight path recovery methods. Applications in geological mapping, identification of structural features and altered zones.

A2. Marine Geophysics:

A3. Geophysical Signal Processing:

A4. Remote Sensing and Geohydrology:
Fundamental concepts of remote sensing, electromagnetic radiation spectrum, Interaction of electromagnetic energy and its interactions in atmosphere and surface of the earth, elements of photographic systems, reflectance and emittance, false color composites, remote sensing platforms, flight planning, geosynchronous and sun synchronous orbits, sensors, resolution, parallax and vertical exaggeration, relief displacement, mosaic, aerial photo interpretation and geological application. Fundamentals of photogrammetry, satellite remote sensing, multi-spectral scanners, thermal scanners, microwave remote sensing, fundamental of image processing and interpretation for geological applications. Types of water bearing formations, porosity, permeability, storage coefficient, specific storage, specific retention, specific yield, Different types of aquifers, vertical distribution of ground water, General flow equation; steady and unsteady flow of ground water in unconfined and confined aquifers.

PART-B

B1. Solid State Physics and Basic Electronics
Crystalline and amorphous structure of matter; Different crystal systems, Space groups; Methods of
determination of crystal structure; X-ray diffraction, Scanning and transmission electron
microscopes; Band theory of solids, conductors, insulators and semiconductors; Thermal properties
of solids, Specific heat: Einstein’s and Debye theory; Magnetism: dia, para and ferro; Elements of
superconductivity; Meissner effect, Josephson junctions and applications; Elementary ideas about
high temperature superconductivity.

Semiconductor devices and circuits: Intrinsic and Extrinsic semiconductors; Devices and structures
(p-n junctions, diodes, transistors, FET, JFET and MOSFET, homo and hetero junction transistors,
thermistors), Device characteristics, Frequency dependence and applications. Opto-electronic
devices (solar cells, photo detectors, LEDs) Operational amplifiers and their applications.

B2. Laser systems
Spontaneous and stimulated emission of radiation. Coherence, Light amplification and relation
between Einstein A and B coefficients. Rate equations for three and four level systems. Lasers: Ruby,
Nd-YAG, CO₂, Dye, Excimer, Semiconductor. Laser cavity modes, Line shape function and full width
at half maximum (FWHM) for natural broadening, collision broadening, Doppler broadening;
Saturation behavior of broadened transitions, Longitudinal and transverse modes. Mode selection,
ABCD matrices and cavity stability criteria for confocal resonators. Quality factor, Expression for
intensity for modes oscillating at random and mode-locked in phase. Methods of Q-switching and
mode locking. Optical fiber waveguides, Fiber characteristics.

B3. Digital electronics, Radar systems, Satellite communications
Digital techniques and applications: Boolean identities, de Morgan’s theorems, Logic gates and truth
tables; Simple logic circuits: registers, counters, comparators and similar circuits. A/D and D/A
converters. Microprocessor: basics and architecture; Microcontroller basics. Combination and
sequential logic circuits, Functional diagram, Timing diagram of read and write cycle, Data transfer
processing, Surveillance radar, Tracking radar, Radar antenna parameters. Fundamentals of
satellite systems, Communication and Orbiting satellites, Satellite frequency bands, Satellite orbit
and inclinations. Earth station technology.

B4. Quantum Mechanics
Wave-particle duality; Wave functions in coordinate and momentum representations; Commutators
and Heisenberg’s uncertainty principle; Schrodinger’s wave equation (time-dependent and time-
independent); Eigenvalue problems: particle in a box, harmonic oscillator, tunneling through a 1-D
barrier; Motion in a central potential; Orbital angular momentum; Addition of angular momentum;
Hydrogen atom; Matrix representation; Dirac’s bra and ket notations; Time-independent
perturbation theory and applications; Variational method; WKB approximation; Time dependent
perturbation theory and Fermi’s Golden Rule; Selection rules; Semi-classical theory of radiation;
Elementary theory of scattering, Phase shifts, Partial waves, Born approximation; Identical particles,
Pauli’s exclusion principle, Spin-statistics connection; Relativistic quantum mechanics: Klein
Gordon and Dirac equations.

Stage-II (Descriptive Type)
Chemistry : Paper-I (Inorganic Chemistry)

1. Inorganic solids:
Defects, non-stoichiometric compounds and solid solutions, atom and ion diffusion, solid
electrolytes. Synthesis of materials, monoxides of 3d-metals, higher oxides, complex oxides
(corundrum, ReO₃, spinel, perovskites), framework structures (phosphates, aluminophosphates,
silicates, zeolites), nitrides and fluorides, chalcogenides, intercalation chemistry, semiconductors,
molecular materials.

2. Chemistry of coordination compounds:
Isomerism, reactivity and stability: Determination of configuration of cis- and trans- isomers by
chemical methods. Labile and inert complexes, substitution reactions on square planar complexes,
trans effect. Stability constants of coordination compounds and their importance in inorganic
analysis.

Structure and bonding: Elementary Crystal Field Theory: splitting of d⁶ configurations in
octahedral, square planar and tetrahedral fields, crystal field stabilization energy, pairing energy.
Jahn-Teller distortion. Metal-ligand bonding, sigma and pi bonding in octahedral complexes and their effects on the oxidation states of transition metals. Orbital and spin magnetic moments, spin only moments and their correlation with effective magnetic moments, d-d transitions; LS coupling, spectroscopic ground states, selection rules for electronic spectral transitions; spectrochemical series of ligands, charge transfer spectra.

3. Acid base titrations:
Titration curves for strong acid-strong base, weak acid-strong base and weak base-strong acid titrations, polyprotic acids, poly-equivalent bases, determining the equivalence point: theory of acid-base indicators, pH change range of indicator, selection of proper indicator. Principles used in estimation of mixtures of NaHCO₃ and Na₂CO₃ (by acidimetry).

4. Gravimetric Analysis:
General principles: Solubility, solubility product and common ion effect, effect of temperature on the solubility; Salt hydrolysis, hydrolysis constant, degree of hydrolysis. Stoichiometry, calculation of results from gravimetric data. Properties of precipitates. Nucleation and crystal growth, factors influencing completion of precipitation. Co-precipitation and post-precipitation, purification and washing of precipitates. Precipitation from homogeneous solution. A few common gravimetric estimations: chloride as silver chloride, sulphate as barium sulphate, aluminium as oxinate and nickel as dimethyl glyoximate.

5. Redox Titrations:

6. Complexometric titrations:
Complex formation reactions, stability of complexes, stepwise formation constants, chelating agents. EDTA: acidic properties, complexes with metal ions, equilibrium calculations involving EDTA, conditional formation constants, derivation of EDTA titration curves, effect of other complexing agents, factors affecting the shape of titration curves: indicators for EDTA titrations, titration methods employing EDTA: direct, back and displacement titrations, indirect determinations, titration of mixtures, selectivity, masking and demasking agents. Typical applications of EDTA titrations: hardness of water, magnesium and aluminium in antacids, magnesium, manganese and zinc in a mixture, titrations involving unidentate ligands: titration of chloride with Hg₂⁺ and cyanide with Ag⁺.

7. Organometallic compounds:
18-electron rule and its applications to carbonyls and nature of bonding involved therein. Simple examples of metal-metal bonded compounds and metal clusters. Wilkinson’s catalyst.

8. Nuclear chemistry:

9. Chemistry of d- and f-block elements:
**d-block elements:** General comparison of 3d, 4d and 5d elements in terms of electronic configuration, elemental forms, metallic nature, atomization energy, oxidation states, redox properties, coordination chemistry, spectral and magnetic properties.

**f-block elements:** Electronic configuration, ionization enthalpies, oxidation states, variation in atomic and ionic (3+) radii, magnetic and spectral properties of lanthanides, separation of lanthanides (by ion-exchange method).

**Stage-II (Descriptive Type)**

**Chemistry : Paper-II (Physical Chemistry)**

1. Kinetic theory and the gaseous state:
Real gases, Deviation of gases from ideal behaviour; compressibility factor; van der Waals equation of state and its characteristic features. Existence of critical state. Critical constants in terms of van
under Waals constants. Law of corresponding states and significance of second virial coefficient. Boyle temperature.

2. **Solids:** Nature of solid state. Band theory of solids: Qualitative idea of band theory, conducting, semiconducting and insulating properties.

Law of constancy of angles, concept of unit cell, different crystal systems, Bravais lattices, law of rational indices, Miller indices, symmetry elements in crystals. X-ray diffraction, Bragg’s law.

3. **Chemical thermodynamics and chemical equilibrium:**

   Chemical potential in terms of Gibbs energy and other thermodynamic state functions and its variation with temperature and pressure. Gibbs-Duhem equation; fugacity of gases and fugacity coefficient. Thermodynamic conditions for equilibrium, degree of advancement. vant Hoff’s reaction isotherm. Equilibrium constant and standard Gibbs energy change. Definitions of $K_p$, $K_c$ and $K_x$; vant Hoff’s reaction isobar and isochore. Activity and activity coefficients of electrolytes / ions in solution. Debye-Hückel limiting law.

4. **Chemical kinetics and catalysis:**


   Physisorption and chemisorption, adsorption isotherms, Freundlich and Langmuir adsorption isotherms, BET equation, surface area determination; colloids, electrical double layer and colloid stability, electrokinetic phenomenon. Elementary ideas about soaps and detergents, micelles, emulsions.

5. **Electrochemistry:**


   Basic principles of pH metric and potentiometric titrations, determination of equivalence point and pK_a values.

6. **Quantum chemistry:**


7. **Basic principles and applications of spectroscopy:**

   Electromagnetic radiation, interaction with atoms and molecules and quantization of different forms of energies. Units of frequency, wavelength and wavenumber. Condition of resonance and energy of absorption for various types of spectra; origin of atomic spectra, spectrum of hydrogen atom.

   **Rotational spectroscopy of diatomic molecules:** Rigid rotor model, selection rules, spectrum, characteristic features of spectral lines. Determination of bond length, effect of isotopic substitution.

   **Vibrational spectroscopy of diatomic molecules:** Simple Harmonic Oscillator model, selection rules and vibration spectra. Molecular vibrations, factors influencing vibrational frequencies. Overtones, anharmonicity, normal mode analysis of polyatomic molecules.

   **Raman Effect:** Characteristic features and conditions of Raman activity with suitable illustrations. Rotational and vibrational Raman spectra.

8. **Photochemistry:**


*Stage-II (Descriptive Type)*
A1. Errors in quantitative analysis:
Accuracy and precision, sensitivity, specific standard deviation in analysis, classification of errors
and their minimization, significant figures, criteria for rejection of data, Q-test, t-test, and F-test,
control chart, sampling methods, sampling errors, standard reference materials, statistical data
treatment.

A2. Separation Methods:
Chromatographic analysis: Basic principles of chromatography (partition, adsorption and ion
exchange), column chromatography, plate concept, plate height (HETP), normal phase and reversed
phase concept, thin layer chromatography, frontal analysis, principles of High Performance Liquid
Chromatography (HPLC) and Gas Liquid Chromatography (GLC), and Ion-exchange chromatography.
Solvent extraction: Classification, principle and efficiency of the technique, mechanism of
extraction, extraction by solvation and chelation, qualitative and quantitative aspects of solvent
extraction, extraction of metal ions from aqueous solutions.

A3. Spectroscopic methods of analysis:
Lambert-Beer's Law and its limitations.
UV-Visible Spectroscopy: Basic principles of UV-Vis spectrophotometer, Instrumentation consisting
of source, monochromator, grating and detector, spectrophotometric determinations (estimation of
metal ions from aqueous solutions, determination of composition of metal complexes using Job's
method of continuous variation and mole ratio method).
Infra-red Spectrometry: Basic principles of instrumentation (choice of source, monochromator and
detector) for single and double beam instruments, sampling techniques.
Flame atomic absorption and emission spectrometry: Basic principles of instrumentation (choice of
source, monochromator, detector, choice of flame and burner design), techniques of atomization
and sample introduction, method of background correction, sources of chemical interferences and
methods of removal, techniques for the quantitative estimation of trace level metal ions. Basic
principles and theory of AAS. Three different modes of AAS - Flame-AAS, VG-AAS, and GF-AAS.
Single beam and double beam AAS. Function of Hollow Cathode Lamp (HCL) and Electrode
Discharge Lamp (EDL). Different types of detectors used in AAS. Qualitative and quantitative
analysis.

A4. Thermal methods of analysis:
Theory of thermogravimetry (TG), basic principle of instrumentation, techniques for quantitative
analysis of Ca and Mg compounds.

A5. X-ray methods of Analysis:
Introduction, theory of X-ray generation, X-ray spectroscopy, X-ray diffraction and X-ray
fluorescence methods, instrumentation and applications. Qualitative and quantitative
measurements. Powder diffraction method.

A6. Inductively coupled plasma spectroscopy:
Theory and principles, plasma generation, utility of peristaltic pump, sampler–skimmer systems, ion
lens, quadrupole mass analyzer, dynode / solid state detector, different types of interferences-
spectroscopic and non-spectroscopic interferences, isobaric and molecular interferences,
applications.

A7. Analysis of geological materials:
Analysis of minerals and ores- estimation of (i) CaCO₃, MgCO₃ in dolomite (ii) Fe₂O₃, Al₂O₃, and TiO₂
in bauxite (iii) MnO and MnO₂ in pyrolusite. Analysis of metals and alloys: (i) Cu and Zn in brass (ii)
Cu, Zn, Fe, Mn, Al and Ni in bronze (iii) Cr, Mn, Ni, and P in steel (iv) Pb, Sb, Sn in ‘type metal’.
Introduction to petroleum: constituents and petroleum fractionation. Analysis of petroleum
products: specific gravity, viscosity, Doctor test, aniline point, colour determination, cloud point,
pour point. Determination of water, neutralization value (acid and base numbers), ash content,
Determination of lead in petroleum.
Types of coal and coke, composition, preparation of sample for proximate and ultimate analysis,
calorific value by bomb calorimetry.

PART B (Organic chemistry)
B1. Unstable, uncharged intermediates:
Structure and reactivity of carbenes and nitrenes and their rearrangements (Reimer-Tiemann, Hoffman, Curtius, Lossen, and Schimdt.).

B2. Addition reactions:
**Addition to C-C multiple bonds:** Mechanism of addition involving electrophiles, nucleophiles and free radicals (polymerization reactions of alkenes and substituted alkenes), Ziegler-Natta catalyst for polymerization, polyurethane, and conducting polymers; addition to conjugated systems (Diels-Alder reaction), orientation and reactivity (on simple cis- and trans- alkenes).

**Addition to carbon-heteroatom multiple bonds:** Addition to C=O double bond, structure and reactivity, hydration, addition of ROH, RSH, CN-, bisulphite, amine derivatives, hydride ions.

B3: Reactions at the carbonyl group:
Cannizzaro, Aldol, Perkin, Claisen ester, benzoin, benzil-benzilic acid rearrangement, Mannich, Dieckmann, Michael, Strobe, Darzen, Wittig, Doebner, Knoevenagel, Reformatsky reactions.

B4. Oxidation and Reduction:
Reduction of C=C, Meerwein-Pondorf reaction, Wolff-Kishner and Birch reduction.
Oxidation of C=C, hydration, hydroxylation, hydroboration, ozonolysis, epoxidation, Sharpless epoxidation.

B5. Electrocyclic Reactions:
Molecular orbital symmetry, frontier orbitals of ethylene, 1,3-butadiene, 1,3,5-hexatriene, allyl system, FMO approach, pericyclic reactions, Woodward-Hoffman correlation diagram method and perturbation molecular orbital (PMO) approach for the explanation of pericyclic reactions under thermal and photochemical conditions. Simple cases of Norrish type-I and type-II reactions. Conrotatory and disrotatory motions of (4n) and (4n+2) polyenes with emphasis on [2+2] and [4+2] cycloadditions, sigmatropic rearrangements- shift of H and carbon moieties, Claisen, Cope, Sommerlet-Hauser rearrangement.

B6. Spectroscopic methods of analysis:

**Ultraviolet spectroscopy:** Chromophores, auxochromes. Electronic transitions ($\sigma-\sigma^*$, $n-\sigma^*$, $\pi-\pi^*$ and $n-\pi^*$), relative positions of $\lambda_{\text{max}}$ considering conjugative effect, steric effect, solvent effect, red shift (bathochromic shift), blue shift (hypsochromic shift), hyperchromic effect, hypochromic effect (typical examples). Woodward rules. Applications of UV spectroscopy to conjugated dienes, trienes, unsaturated carbonyl compounds and aromatic compounds.

**Nuclear Magnetic Resonance Spectrometry:** (Proton and Carbon-13 NMR) Nuclear spin, NMR active nuclei, principle of proton magnetic resonance, equivalent and non-equivalent protons. Measurement of spectra, the chemical shift, shielding / deshielding of protons, upfield and downfield shifts, intensity of NMR signals and integration factors affecting the chemical shifts: spin-spin coupling to $^{13}$C $^1$H-$^1$H first order coupling: some simple $^1$H-$^1$H splitting patterns: the magnitude of $^1$H-$^1$H coupling constants, diamagnetic anisotropy.

**Mass spectrometry:** Basic Principles, the mass spectrometer, isotope abundances; the molecular ion, metastable ions. McLafferty rearrangement.